

Molar Mass

Worksheet-Answer Key

Name _____ Date _____

Exercise 1. How much does one mole weigh (molar mass/molecular weight) for the following? Show work: which numbers are added/multiplied, how many decimal places are used from the periodic table, and your result.

Na 23g

22.99 → 23 g

NaF 42g

23 + 19 = 42 g

Zn 65.4 g

65.38 → 65.4 g

CO 28 g

12 + 16 = 28 g

solid iron 55.9 g

Fe 55.85 → 55.9 g

CO₂ 44 g

12 + 2×16 = 12 + 32 = 44 g

Ca²⁺ (aq) 40 g

same as Ca 40.08 → 40 g

CH₄ 16 g

12 + 4×1 = 12 + 4 = 16

CaCO₃ 100 g

40 + 12 + 3×16 = 100g

C₄H₁₀ 58 g

4×12 + 10×1 = 58g

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$$2 \times 19 = 38 \text{ g}$$



$$2 \times 12 + 2 \times 1 = 24 + 2 = 26 \text{ g}$$



$$2 \times 14 = 28 \text{ g}$$



$$2 \times 12 + 4 \times 1 = 24 + 4 = 28 \text{ g}$$



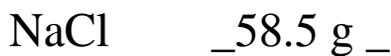
$$2 \times 16 = 32 \text{ g}$$



$$2 \times 55.9 + 3 \times 16 = 111.8 + 48 = 159.8 \text{ g } 159.8 \text{ g}$$



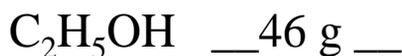
$$2 \times 35.5 = 71 \text{g}$$



$$23 + 35.5 = 58.5 \text{g}$$



$$39 + 35.5 + 3 \times 16 = 122.5 \text{g}$$



$$2 \times 12 + 6 \times 1 + 16 = 46 \text{g}$$